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1) Transition metals. Iron + bromine iron (III) bromide. $2\text{Fe (s)} + 3\text{Br}_2 \text{ (l)} \rightarrow 2\text{FeBr}_3 \text{ (s)}$ Iron + chlorine iron (III) chloride. $2\text{Fe (s)} + 3\text{Cl}_2 \text{ (g)} \rightarrow 2\text{FeCl}_3 \text{ (s)}$ All of the compounds with metals are ionic salts which form a giant structure. They are called metal halides because they are formed from a metal and a halogen.

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Similarly potassium and bromine form potassium bromide KBr, or lithium and iodine form lithium iodide LiI Again note the group formula pattern Hydrochloric acid reacts with metals above hydrogen in the reactivity series

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QUESTION - CHEMISTRY FORM -2 CHEMICAL FAMILIES 1. Study the information in the table below and answer the questions that follow: Element Atomic radius (nm) Ionic radius (nm) W 0.114 0.195 X 0.072 0.136 Y 0.133 0.216 Z 0.099 0.181 (a) Would these form part of a metallic or a non-metallic group? Explain

QUESTION - CHEMISTRY FORM -2

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