

# Chapter 18 Review Chemical Equilibrium Answers Section 3

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CHAPTER 18: Chemical Equilibrium I. a.

Write the chemical equilibrium expression for the following equations.

Include the value of  $K$ . (a)  $(g) \rightleftharpoons (g)$   $K=0.1$

(b)  $(g) + O_2(g) \rightleftharpoons (g)$   $K = 1.8 \times 10^{-2}$  Does the reaction favor products or reactants?

(Will there be mostly products or reactants when it reaches 2. a. Compare

the rates of forward and reverse

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reactions when equilibrium has been reached. b.

### **North St. Paul-Maplewood Oakdale / Overview**

Modern Chemistry 145 Chemical  
Equilibrium CHAPTER 18 REVIEW  
Chemical Equilibrium SECTION 2 SHORT  
ANSWER Answer the following questions  
in the space provided. 1. \_\_\_\_ Raising  
the temperature of any equilibrium  
system always (a) favors the forward  
reaction. (b) favors the reverse reaction.  
(c) favors the exothermic reaction. (d)  
favors the endothermic reaction. 2.  
Consider the following equilibrium  
equation:

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svlsarah. chapter 18 & 19 chemistry;  
Chemical equilibrium. Equilibrium.  
reversible reaction. chemical  
equilibrium. law of chemical equilibrium.  
A state of balance. can occur in both the  
forward and reverse directions. state of

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balance in which forward and reverse reactions take p....

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What does it mean when a reaction reaches a chemical equilibrium When the forward and reverse reactions balance each other out because they take place at the same rate What

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happens to the equilibrium when products are removed or added It shifts to the side where the products originated from

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8.4 Chapter Summary (ESCNV).  
Presentation: 27V5. A reaction is reversible when reactants can react to form products, and products can react to form the reactants again.. A reaction is in chemical equilibrium when the rate of the forward reaction equals the rate of

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the reverse reaction.. In an open system energy and matter can enter and leave the system. In a closed system energy can enter and ...

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1. In a bottle of unopened cola, the CO<sub>2</sub> gas dissolved in the liquid is in equilibrium with the CO<sub>2</sub> gas above the liquid. The dissolved gas reacts with water molecules in the cola to form carbonic acid, which also dissociates into carbon dioxide and water. Which



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chemical equation(s) best describe this equilibrium system? a.  $\text{CO}_2(\text{g}) \rightleftharpoons \text{CO}_2(\text{l})$   
b.

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## **Chapter 18 Review Chemical Equilibrium Section 3 Answers**

The balanced chemical equation for an equilibrium system helps establish the expression for the equilibrium constant. The data in Table 18-1 show that the validity of this expression is confirmed when the actual values of the equilibrium concentrations of reactants

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and products are determined experimentally. The values of  $K_c$  are calculated from these concentrations.

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