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Section 2.3 Chemical Properties (pages 54-58) This section discusses chemical properties and describes clues that may show that a chemical change has taken place. Reading Strategy (page 54) Relating Text and Visuals As you read, complete the table by finding examples of the clues for recognizing chemical changes in Figures 19 and 20.

Chapter 2 Properties of Matter Section 2.3 Chemical Properties

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CHEMICAL . First the teeth just make a change in size, and then

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the enzymes make a chemical change producing new molecules. Our temperature goes up while we digest from the energy released. 10. In a fireworks show, the fireworks explode giving off heat and light. CHEMICAL. Substances giving off heat and light, themselves, is always chemical.

ANSWERS Physical/Chemical Properties/Change - Google Docs

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13. Scientists and engineers review the physical and chemical properties of materials to determine their usefulness. (1/1) 14. Ceramics and polymers vary widely in their. chemical composition. The chemical composition determines their physical and chemical properties. (3/1) 15. A scientist works to learn more scientific

Teacher Guide & Answers - Glencoe

The change of one type of matter into another type (or the inability to change) is a chemical property. Examples of chemical properties include flammability, toxicity, acidity, reactivity (many types), and heat of combustion. Iron, for example, combines with oxygen in the presence of water to form rust; chromium does not oxidize . Nitroglycerin is very dangerous because it explodes easily; neon poses almost no hazard because it is very unreactive.

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1.3 Physical and Chemical Properties - Chemistry

Chemical properties describe matter based on its ability to change into new matter. One chemical property of matter is reactivity. Reactivity is the ability of a substance to change into a new substance. One kind of reactivity is flammability.

1 SECTION 3 Chemical Properties

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Chemical Properties. A chemical property describes the ability of a substance to undergo a specific chemical change. A chemical property of iron is that it is capable of combining with oxygen to form iron oxide, the chemical name of rust. The more general term for rusting and other similar processes is corrosion.

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2.13: Chemical Properties and Chemical Reactions ...

Interactive Reader and Study Guide 52 Properties of Matter
SECTION3 Chemical Properties Properties of Matter Name Class
Date CHAPTER 3 California Science Standards 8.5.a, 8.5.c, 8.5.d
STUDY TIP Compare Make a table with two columns: Chemical
property and Physical property. List the chemical and physical
properties that are discussed in this ...

CHAPTER SECTION 3 Chemical Properties

The change of one type of matter into another type (or the inability to change) is a chemical property. Examples of chemical properties include flammability, toxicity, acidity, reactivity (many types), and heat of combustion. Iron, for example, combines with oxygen in the presence of water to form rust; chromium does not oxidize (Figure 2).

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Physical and Chemical Properties | Chemistry

Key Concept A chemical property describes the ability of a substance to change into a new substance. What You Will Learn.

- Examples of chemical properties are reactivity and flammability.
- A chemical change is the process by which a substance changes into a new substance.
- Chemical changes usually liberate or absorb heat.

Section 3 Chemical Properties - Travellin

Chemical properties describe matter based on its ability to change into new matter that has different properties. How do chemical properties differ from physical properties? Physical properties do not change matter. Burning wood: there are two new substances with different

Chapter 2 Section 3; Chemical Properties by Barbara Knight

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- (4 characteristic properties) 1. Act as a unit 2. Smallest unit of a substance that has the same properties of the substance 3. Some molecules made of different elements are also compounds 4. Some molecules are made of atoms of the same element

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